

Chem 14A Su2021

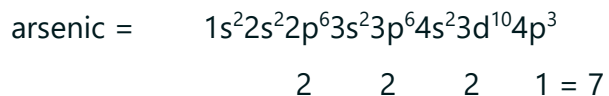
Week 2 Quiz

How many of arsenic's electrons (in its ground state configuration) have $m_l = +1$?

Enter only a number for your answer.

2 electrons per orbital max.

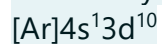
$m_l = +1$ refers to only one p and d orbital per energy level (shell) in arsenic



How many of argon's electrons (in its ground state configuration) have $m_s = +1/2$?

Exactly half. = 9

How many of copper's electrons (in its ground state configuration) are unpaired?



So only 1 electron is unpaired

How many different values of the quantum number l are possible when $n = 4$?

4s $l = 0$

4p $l = 1$

4d $l = 2$

4f $l = 3$

Total of 4 different l values.

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